Synthesis and characterization of ferric tannate as a novel porous adsorpive-catalyst for nitrogen removal from wastewater

Ruina Zhang, Lin Li and Junxin Liu*

Ferric tannate has unique adsorption capacities for NH$_4^+$-N and NO$_2^-$-N simultaneously. So far, no normative method is available for synthesizing ferric tannate. In this work, an optimized method for synthesizing ferric tannate by using tannic acid and ferric chloride is established under the conditions of room temperature and neutral pH. The optimal molar ratio of tannic acid and ferric chloride were determined to be in the range of 1:20 to 1:25 based on the yield and stability of the ferric tannate composite. Scanning electron microscopy, the Brunauer–Emmett–Teller method, X-ray diffraction, and Fourier transform infrared spectrometry were used to characterize the ferric tannate composite. The results showed that the ferric tannate has a rough and porous surface, a large surface area, and an amorphous structure. Experiments of NH$_4$+$^+-$N and NO$_2^-$-N adsorption reflected that the ferric tannate composite has a unique capacity for adsorption-catalyzed conversion of NH$_4$+$^+-$N and NO$_2^-$-N to N$_2$. When the mass ratios of ferric tannate/NH$_4$+$^+-$N and ferric tannate/NO$_2^-$-N were both 200, the removal of NH$_4$+$^+-$N and NO$_2^-$-N was 98.1% and 96.2%, respectively, after 3.0 hours of reaction. The conversion rate of N$_2$ increased to 87.1%. Therefore, ferric tannate can be used as a potential material for removing nitrogen from wastewater in the future.

**Introduction**

Nitrogen in wastewater is acknowledged as one of the critical causes of severe eutrophication of receiving water bodies. In order to satisfy the strict sewage discharge standards, many technologies for wastewater nitrogen removal have been extensively developed, such as biological nitrogen removal, chemical oxidation, adsorption, and ion exchange. Currently, biological nitrogen removal through the combination of aerobic nitrification and anoxic denitrification is widely applied in wastewater treatment plants. However, this approach requires a large amount of oxygen for nitrification, and a sufficient amount of organic substrates to maintain denitrification, which results in high energy consumption. Another biotechnology, anaerobic ammonium oxidation (Anammox), is recognized as a promising alternative for nitrogen removal. In the Anammox process, NO$_2^-$-N is converted to nitrogen gas (N$_2$) with NH$_4$+$^+$-N as the electron donor; this process consumes less energy and does not require additional carbon substrate as the electron donor. However, Anammox bacteria exhibit an extremely slow growth rate, limiting its wide practical application. The chemical approach for nitrogen removal gives a greater reaction rate and a higher removal efficiency, but this approach requires additional oxidants or reductants, thereby causing high costs and complicating operational procedures.

Adsorption and ion exchange approaches, such as using activated carbon and zeolites, can quickly remove NH$_4$+$^+$-N, NO$_2^-N$, and NO$_3^-$-N from wastewater. However, the adsorbent and ion exchange resins must be regenerated after the saturation of adsorbed nitrogen, and the regenerated solution needs further disposal, leading to higher environmental and economical costs. Thus, adsorption would be an attractive alternative approach for wastewater nitrogen removal if the adsorbent resin can be regenerated in situ.

Theoretically, the redox reaction for N$_2$ production occurs between NH$_4$+$^+$-N and NO$_2^-N$ based on their standard redox potential, but the reaction rate depends on the reaction conditions significantly. Chemical kinetics indicate that the reaction rate can be accelerated through increasing the concentrations of reactants. Adsorption can transform reactants from liquids into solids, causing the reactants to be concentrated on the surface of adsorbents. A very recent patent implied that the presence of ferric tannate (a kind of magnetic nanoparticle) is capable of accelerating this reaction at room temperature. Furthermore, a highly cross-linked network of ferric tannate was found to exhibit a rough and porous surface, capable of absorbing both negative and cationic ions due to its ferric ions and negative oxygen ions. In addition, N$_2$H$_4$ and NH$_2$OH are intermediate products in the reaction of NO$_2^-N$ and NH$_4$+$^+$-N while ferric ions can convert...
$\text{NH}_4\text{H}_4$ and $\text{NH}_2\text{OH}$ to $\text{N}_2$. Thus, ferric tannate may be used as a novel solution for nitrogen removal. Even though many previous studies have illustrated the development and applications of magnetic nanoparticles in nitrogen removal, no literature has clearly demonstrated that these magnetic nanoparticles, including ferric tannate, could efficiently eliminate \(\text{NH}_4^+\)-N and \(\text{NO}_2^\text{-N}\) simultaneously, as well as accelerate the redox reaction between these two substances. Additionally, no approach has been developed to date for synthesizing ferric tannate that can be efficiently used in wastewater treatment.

Therefore, this study aims to develop an optimized approach for synthesizing ferric tannate, which can accelerate the reaction between \(\text{NH}_4^+\)-N and \(\text{NO}_2^\text{-N}\). Tannic acid and ferric chloride were used as the raw materials to synthesize ferric tannate under the combined conditions of room temperature and neutral pH. Moreover, Fourier transform infrared spectrometry (FTIR) and the Brunauer–Emmett–Teller (BET) method were used to analyze the functional groups and surface morphology of the produced ferric tannate. This study also investigated the potential nitrogen removal performance of ferric tannate in wastewater treatment. The objective of this study is to synthesize a material that can accelerate the reaction between \(\text{NO}_2^\text{-N}\) and \(\text{NH}_4^+\)-N.

## Materials and methods

### Materials

Tannic acid and ferric chloride (FeCl₃) were used to synthesize ferric tannate. Tannic acid (C₇₆H₅₂O₄₆, molar mass: 1701.18 g mol⁻¹) was purchased from Sigma-Aldrich Co., USA, with a pure grade of ≈99.5%. FeCl₃, NH₄Cl, and NaNO₂ were purchased from Sinopharm Chemical Reagent Co., China, with a pure grade of ≈99.5%. NaHCO₃ was purchased from Tianjin Jinke Fine Chemical Research Institute, China, with a pure grade of ≈99.5%. The solutions of tannic acid, FeCl₃, and NaHCO₃ were prepared with distilled water at concentrations of 0.1, 1, and 0.65 mol L⁻¹, respectively.

### Synthesis of ferric tannate

Tannic acid is an abundant polyphenolic compound. Its ortho-phenolic hydroxyl structure can cause complexation reactions with ferrous and ferric ions, which form a hybrid with metal-organic frameworks. The chemical structures of tannic acid and ferric tannate are shown in Fig. 1. In view of the complexity of the chemical structure of tannic acid, six molar ratios of tannic acid and FeCl₃ (1 : 10, 1 : 15, 1 : 20, 1 : 25, 1 : 30, and 1 : 40) were chosen to investigate the influence of the varying proportions on ferric tannate yield. The synthesis of ferric tannate was performed in a conical flask with a volume of 150 mL. An amount of 10 mL of 0.1 mol L⁻¹ tannic acid solution was mixed with various volumes of 1 mol L⁻¹ FeCl₃ solution. The pH of the mixture was adjusted to 7 by using 0.65 mol L⁻¹ NaHCO₃ solution. The reaction products were separated via sedimentation, and the resultant precipitate was washed with deionized water four times, centrifuged (3000 rpm, 2 min), and dried in a freezer dryer at −50 °C.

### Characterization of ferric tannate

The ferric tannate yield (\(y_c\)) was calculated using the following equation (eqn (1)):

\[
y_c = \frac{m_b}{(m_t + m_{ic})} \times 100\%
\]

where \(m_b\) is the mass of the ferric tannate (mg), and \(m_t\) and \(m_{ic}\) are the masses of tannic acid and ferric chloride (mg), respectively.

To determine the stability of the ferric tannate yield, 0.2 g of the ferric tannate composition, which was prepared with different ratios as mentioned earlier, was placed into six shake flasks. Then, each shake flask was charged with 200 mL of deionized water. After 1 h of continuous stirring, the mixture in each shake flask was filtered through a 0.45 μm pore size filter. The filtrate was measured with a UV-vis spectrophotometer (UV-1700, Hitachi Co., Japan) at a 200 nm to 500 nm wavelength.

Field emission scanning electron microscopy (FE-SEM), which was coupled with an energy dispersive X-ray spectrometry (EDS) detector (SU-8020, Hitachi Co., Japan), was employed to observe the surface physical morphology and major elements of ferric tannate.

An automatic micropore and mesopore analyzer (ASAP 2020 HD88, USA) was operated to analyze the specific surface area and pore size distribution of ferric tannate. Moreover, the specific surface area and pore size distribution of the ferric tannate were determined with a N₂ adsorption isotherm at 77 K by using a Micromeritics ASAP 2010 accelerated surface area analyzer and the BET method.

Fourier transform infrared spectrometry (FTIR, Tensor 27, Germany) was applied to measure the functional groups of tannic acid and ferric tannate with the potassium bromide pellet method in the range of 400 cm⁻¹ to 4000 cm⁻¹.

X-ray diffraction (XRD, XPert PRO MPD, PANalytical, Holland) was used to analyze the purity and crystallinity of ferric tannate, ferric chloride, and tannic acid with Cu-Kα radiation (\(\lambda = 0.15418\) nm) in the 2θ range of 10° to 80°. The XRD data were collected in a scan mode with a step length of 0.5° with a scanning rate of 6° min⁻¹. The accelerating voltage and the applied current were 45 kV and 45 mA, respectively.

### Determination of the nitrogen removal performance

Several batch tests were conducted in stoppered conical flasks to estimate the nitrogen removal performance of the synthesized ferric tannate. Stock solutions were prepared by dissolving NH₄Cl and/or NaNO₂ in deionized water. Each time, 0.1 g ferric tannate and 50 mL stock solution were mixed in a flask, which was then shaken with a shaker at 200 rpm at room temperature. In the procedure of nitrogen removal by using ferric tannate, the pH was 7, which was adjusted with HCl or NaHCO₃.

In the adsorption isotherm studies, stock solutions with different initial concentrations of \(\text{NH}_4^+\)-N or \(\text{NO}_2^\text{-N}\) were added into 100 mL conical flasks, and the equilibrium time was set to 2 h. The samples were separated via filtration.

In the experiments, the concentrations of both \(\text{NH}_4^+\)-N and \(\text{NO}_2^\text{-N}\) in the reaction solutions were 50 mg L⁻¹. 100 mL stock
solution and 1.0 g ferric tannate were added into a 150 mL stoppered conical flask. Then, the flask was shaken at 200 rpm for 3.0 h at room temperature, and 500 μL water samples and 50 μL gas samples were collected every 20 min. In the study, the data of NH$_4^+$ and NO$_2^-$ adsorption on the ferric tannate at room temperature were obtained through three groups of parallel tests.

The concentrations of NH$_4^+$-N and NO$_2^-$-N were measured according to standard methods.$^{31}$ The Agilent 6890N Gas Chromatograph equipped with a TCD was used to analyse the content of N$_2$.

The adsorption amount of NH$_4^+$-N or NO$_2^-$-N on the ferric tannate at equilibrium, $q_e$ (mgN g$^{-1}$), was calculated with the following equations (eqn (2) and (3)):

$$q_e = \frac{(C_i - C_f)V}{m}$$  \hspace{1cm} (2)
Removal = \left( \frac{C_i - C_e}{C_i} \right) \times 100\% \quad (3)

where \( q_e \) is the equilibrium amount of adsorption on the adsorbent surface (mg g\(^{-1}\)), and \( C_i \) and \( C_e \) are the initial and equilibrium concentrations of NH\(_4\)^+-N or NO\(_2\)^--N (mg N L\(^{-1}\)) in the solution, respectively. \( V \) is the volume of the solution (L), and \( m \) is the mass of the ferric tannate (g).

The Langmuir adsorption equation and the Freundlich adsorption equation can be linearised according to the equation in ref. 32.

Due to the inherent bias resulting from linearisation of the isotherm model, error analysis was employed as a criterion for the fitting quality. This statistical analysis is based on the sum of the squares of the differences between the experimental and the Langmuir model and Freundlich model calculated data. The error was calculated using the following equations (eqn (4) and (5)):

Relative error = \frac{\sum_{i=1}^{n} |q_i - q_{\text{sim}}|}{n} \times 100\% \quad (4)

Standard deviation = \sqrt{\frac{1}{n} \sum_{i=1}^{n} \left( \frac{q_i - q_{\text{sim}}}{q_i} \right)^2 - \left( \frac{1}{n} \sum_{i=1}^{n} \frac{q_i - q_{\text{sim}}}{q_i} \right)^2} \quad (5)

where \( q_i \) is the equilibrium capacity of the adsorbent obtained from the experiment (mg g\(^{-1}\)), and \( q_{\text{sim}} \) is the simulated adsorption capacity which is calculated with the Langmuir equation and the Freundlich equation (mg g\(^{-1}\)).

A low value of the relative error or standard deviation indicates that the experimental data fit better to the value from the model. In order to confirm the best-fit isotherms and kinetic models for the adsorption system, there is a need to analyse the data set using the Nash–Sutcliffe efficiency (NSE).\(^{33,34}\) As a dimensionless goodness-of-fit indicator, the values of the NSE could be widely used and a potentially reliable statistic for assessing the goodness-of-fit of isotherm adsorption models. NSE = 1 indicated a perfect fit, 0.75 \( \leq \) NSE \( \leq \) 0.9 suggested a good fit, and 0.65 \( \leq \) NSE \( \leq \) 0.749 suggested an acceptable fit. NSE is calculated as follows (eqn (6)):
Table 3 The data of error estimation and the NSE

<table>
<thead>
<tr>
<th></th>
<th>Error estimation (%)</th>
<th>NSE</th>
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<tbody>
<tr>
<td></td>
<td>Langmuir</td>
<td>Freundlich</td>
</tr>
<tr>
<td>NH4⁺</td>
<td>NO₃⁻</td>
<td>NH4⁺</td>
</tr>
<tr>
<td>Test1</td>
<td>1.51 ± 1.24</td>
<td>3.50 ± 2.28</td>
</tr>
<tr>
<td>Test2</td>
<td>1.58 ± 0.901</td>
<td>3.51 ± 2.32</td>
</tr>
<tr>
<td>Test3</td>
<td>1.89 ± 1.19</td>
<td>3.85 ± 2.63</td>
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Fig. 8 The removal of NH₄⁺-N and NO₃⁻-N and the conversion rate of N₂ with the addition of ferric tannate.

\[
NSE = 1 - \frac{\sum_{i=1}^{n} (q_i - q_{sim})^2}{\sum_{i=1}^{n} (q_i - \bar{q})^2}
\]

where \(q_i\) is the equilibrium capacity of the adsorbent obtained from the experiment (mg g⁻¹), \(q_{sim}\) is the simulated adsorption capacity (mg g⁻¹), and \(\bar{q}\) is the average adsorption capacity (mg g⁻¹).

Synthesis of ferric tannate

As expected, the reaction of tannic acid with FeCl₃ resulted in the formation of an abundant precipitate of ferric tannate. Table 1 presents a summary of the results of the ferric tannate yield. Increasing the proportion of tannic acid to FeCl₃ from 1 : 10 to 1 : 30 resulted in the gradual increase of ferric tannate yields. The highest yield, 82.6%, appeared in the case of the 1 : 30 ratio. Moreover, the ferric tannate yields are nearly equal in the cases of the 1 : 20, 1 : 25, and 1 : 40 ratios.

The stability of the ferric tannate composition in liquid will be highly important to apply it to nitrogen removal from wastewater. Therefore, the dissolution of the ferric tannate compositions in water was determined. The UV-visible absorbance spectra of the soak water of the ferric tannate compositions are shown in Fig. 2. The results showed that nothing could be dissolved in the cases of the 1 : 15, 1 : 20, and 1 : 25 ratios, but an absorption peak appeared at 302 nm in the case of the 1 : 10 ratio, and at 303 nm in the cases of the 1 : 30 and 1 : 40 ratios, which indicated that some substrates were dissolved into the water.

The preceding results indicated that the molar ratio of tannic acid to ferric ions was responsible for the formation of the ferric tannate complex. With strong reducibility, the phenolic hydroxyl groups in tannic acid can reduce Fe³⁺ to Fe²⁺. In the case of the 1 : 10 ratio, the excess tannic acid can produce a dissolved ferrous tannate, with an absorption peak appearing at 302 nm. In the cases of the 1 : 30 and 1 : 40 ratios, the ferric ion was in excess, which could lead to the formation of Fe(OH)₃. Therefore, the absorption peak at 303 nm is probably from the ferric ion dissolution. Therefore, based on the yield and stability of the ferric tannate compositions, the optimal molar ratio of tannic acid and FeCl₃ to synthesize ferric tannate is between 1 : 20 and 1 : 25.

Characterization of ferric tannate

According to the results of the preceding experiment, the synthesized sample in the case of the 1 : 20 ratio was selected to characterize the ferric tannate composite. FE-SEM images of the ferric tannate composite are illustrated in Fig. 3. The ferric tannate composite had a rough and porous surface structure. The adsorption isotherms of N₂ on the ferric tannate composite were used to deduce the surface characteristics. The results showed that the BET surface area, pore volume, and average pore diameter were 97 m² g⁻¹, 0.073 cm³ g⁻¹, and 66.58 nm, respectively. In general, a large surface area and pore diameter can increase the adsorption capacity of the material.

The SEM-EDS spectrum of the ferric tannate composite is shown in Fig. 4. The composite mainly consists of C, O, and Fe elements, and the percentage content of C, O, and Fe in the ferric tannate composite are 23.06%, 46.83%, and 29.91%, respectively. However, the ratio of C to O from the SEM-EDS spectrum is anomalous compared with the chemical formula of tannic acid (C₇₆H₅₂O₄₆). A possible explanation is that ferric ions partially undergo a hydrolysis reaction at neutral pH, which leads to an increase of O in the ferric tannate composite. However, this finding should be verified with further research.

The XRD patterns of the tannic acid, FeCl₃, and ferric tannate composite are shown in Fig. 5. XRD data indicate that FeCl₃ has many diffraction peaks because of its crystalline structure. The diffraction peaks of the tannic acid and ferric
tannate composite appear in the 2θ = 20° to 30° range, which demonstrates that they have an amorphous structure.

The FTIR spectra of the tannic acid and ferric tannate composite are illustrated in Fig. 6. Two broad adsorption bands in the 3600 cm⁻¹ to 3100 cm⁻¹ range indicating the presence of a phenolic hydroxyl group were observed, and the band at 1320 cm⁻¹ exhibits a characteristic bending vibration of O=C in a phenolic hydroxyl group. The sharp band at 1400 cm⁻¹ in the spectrum of the ferric tannate composite is very pronounced, which can be ascribed to the adsorption peak of a C=O vibration of benzoic acid. This phenomenon can be explained by the chelation and redox reaction of Fe³⁺ with the phenolic hydroxyl group, which results in the coexistence of Fe³⁺ and Fe²⁺ in ferric tannate. The interaction between Fe³⁺ and the phenolic hydroxyl group has an effect on the O=C bond stretching vibration, in which the band is shifted from 1320 cm⁻¹ in the spectrum of the tannic acid to 1338 cm⁻¹ in the spectrum of the ferric tannate composite. During the combination process of Fe³⁺ with the phenolic hydroxyl group, an oxygen anion was formed.

Evaluation of the nitrogen removal performance of the ferric tannate composite

Some batch tests for adsorbing NH₄⁺⁻N and NO₂⁻⁻N were conducted separately to estimate the nitrogen removal performance of the synthesized ferric tannate. The data of NH₄⁺⁻N and NO₂⁻⁻N adsorption on the ferric tannate composites, with only NH₄⁺⁻N or NO₂⁻⁻N in the solution, are shown in Fig. 7. The Langmuir and Freundlich models are usually used to describe the equilibrium isotherm data. The results from linear plots of these isotherms of NH₄⁺⁻N and NO₂⁻⁻N on the ferric tannate composites are listed in Table 2. Compared with the correlation coefficient value of the linear plot of the Freundlich isotherm, that of the Langmuir model was in better agreement with the observed data ($R² = 0.990$ for NH₄⁺⁻N adsorption and $R² = 0.957$ for NO₂⁻⁻N adsorption). The maximum adsorption capacities of NH₄⁺⁻N and NO₂⁻⁻N calculated from the Langmuir equations were 13.6 mg g⁻¹ and 10.2 mg g⁻¹ at room temperature, respectively. The results indicated that the adsorption capacity of ferric tannate toward NH₄⁺⁻N was better than that toward NO₂⁻⁻N.

In the study, the values of the NSE (Table 3) indicated that the Langmuir models of NH₄⁺⁻N and NO₂⁻⁻N can be considered to be very good ($0.9 ≤ NSE ≤ 1$). The Freundlich models of NH₄⁺⁻N and NO₂⁻⁻N can be considered to be good ($0.8 ≤ NSE ≤ 0.9$) and very good ($0.9 ≤ NSE ≤ 1$), respectively.

The adsorption performances of ferric tannate toward NH₄⁺⁻N and NO₂⁻⁻N were different and varied significantly with time. Fig. 8 shows the adsorptive behaviors for NH₄⁺⁻N and/or NO₂⁻⁻N on ferric tannate. The concentrations of NH₄⁺⁻N and NO₂⁻⁻N decreased with time, while the production of N₂ increased over the same time (Fig. 8). When the mass ratios of ferric tannate/ NH₄⁺⁻N and ferric tannate/NO₂⁻⁻N were both 200, the removal efficiencies of NH₄⁺⁻N and NO₂⁻⁻N were 98.1% and 96.2%, respectively, after 3.0 hours of reaction. The conversion rate of N₂ increased to 87.1%. The results showed that after the addition of ferric tannate, the removal of NH₄⁺⁻N and NO₂⁻⁻N increased obviously with N₂ production increasing simultaneously, which indicated that ferric tannate would be beneficial for the reaction between NH₄⁺⁻N and NO₂⁻⁻N, and lead to greater N₂ generation. These results indicated that the ferric tannate had a unique capacity to adsorb anions and cations simultaneously, and an adsorption-catalyzed reaction occurred on the ferric tannate with the adsorption of NH₄⁺⁻N and NO₂⁻⁻N. The adsorption-catalyzed reaction between NH₄⁺⁻N and NO₂⁻⁻N could refresh the adsorption sites of ferric tannate, and the continuous adsorption of NH₄⁺⁻N and NO₂⁻⁻N on the ferric tannate could be promoted. Thus, the removal amount of NH₄⁺⁻N and NO₂⁻⁻N was enhanced. Therefore, the ferric tannate may be used as a promising material to remove nitrogen from wastewater in the future.

For practical use, the dosage of ferric tannate will be adjusted with the concentrations of NH₄⁺⁻N and NO₂⁻⁻N in the wastewater, for the purpose of increasing the amounts of NH₄⁺⁻N and NO₂⁻⁻N converted to N₂.

Conclusions

A novel method to synthesize porous ferric tannate composites was demonstrated by using tannic acid and ferric chloride under the conditions of room temperature and neutral pH. According to the yield and stability of the ferric tannate composite, the optimal molar ratio of tannic acid and FeCl₃ is between 1:20 and 1:25. Combined with the SEM and BET characterization, the characteristics of a porous surface and a large surface area of the ferric tannate composite were confirmed. XRD analysis confirmed that the ferric tannate composite has an amorphous structure. The adsorption test results indicated that the adsorption isotherm of NH₄⁺⁻N and NO₂⁻⁻N on the ferric tannate composite agrees well with that obtained from the Langmuir model. Moreover, the maximum adsorption capacities of NH₄⁺⁻N and NO₂⁻⁻N calculated from the Langmuir equations were 13.6 mg g⁻¹ and 10.2 mg g⁻¹ at room temperature, respectively. Furthermore, ferric tannate has a unique capacity to adsorb NH₄⁺⁻N and NO₂⁻⁻N simultaneously. In addition, an adsorption-catalyzed reaction can occur on the ferric tannate with the adsorption of NH₄⁺⁻N and NO₂⁻⁻N, which enhances the removal efficiency of NH₄⁺⁻N and NO₂⁻⁻N. When the mass ratios of ferric tannate/ NH₄⁺⁻N and ferric tannate/NO₂⁻⁻N were both 200, the removal efficiencies of NH₄⁺⁻N and NO₂⁻⁻N were 98.1% and 96.2%, respectively, after 3.0 hours of reaction. The conversion rate of N₂ increased to 87.1%. Therefore, ferric tannate may be used as a promising material to remove nitrogen from wastewater in the future.

Acknowledgements

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